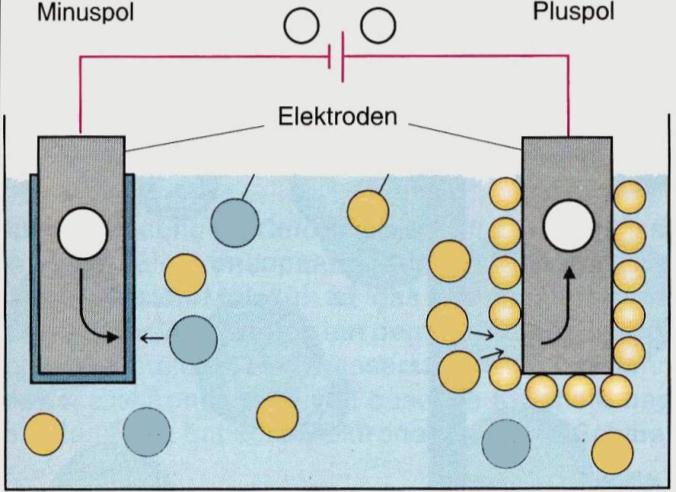
**Electrolysis**

Electrolysis is a method of using an electric current to drive an otherwise non-spontaneous chemical reaction. So basically, it is the inversion of a galvanic element.

Electrolysis is commercially highly important as a stage in the separation of elements from naturally occurring sources such as ores using an electrolytic cell.

**Task**: Zinc bromide is being dissolved in water and an electrical current is applied. Write down oxidation, reduction and all involved particles and substances in the following figure:



**positive pole**

**electrodes**

**negative pole**